

This is all review, you might remember that some review showed up on the last test!

If you want your grade to go up in chemistry you need to score higher than your current % for your grade to go up! If you are scoring lower than your current % on test's then you have to score **a lot higher on labs and hw pkts** for your grade to go up.

Go to Mr C's website <http://sites.google.com/site/mrcleveschemistry/>

How many unit 4 WS's does he have attached to the site? ____ Are you able to open them? ____

Does he have any unit 4 test info up yet? ____

Now the review

An element has an atomic number of 37 and an atomic mass of 86.

Write the electron notation for this element _____

What is the symbol for this element _____

Is this element a metal/non metal/metalloid ? (circle one)

What state would the metal be at room temp? _____

How many valance electrons does this element have? _____

As an Ion what would it's charge be? _____

The outermost electrons of the element are found in what energy level? _____

What is the name of the chemical family that this element belongs to ? _____

If this element combined with carbonate what would it's chemical formula be?

If we had 5.0 moles of the chemical whose formula you just wrote down, how many molecules would that be? (show set up)

What would the mass of that many molecules be? (show set up)

46.0g of sodium metal reacts with 54.0g of water to produce sodium hydroxide and hydrogen gas.

Write a balanced equation

Calculate how many grams of sodium hydroxide would be produced by this reaction.
(careful this is a limiting reagent problem)

Which reactant is the limiting reagent? How do you know?

If in a lab situation you produced 76.5 grams of sodium hydroxide, calculate the percent yield.

At STP (you'll learn this soon) 1 mole of a gas occupies 22.4 Liters. Based on this how many liters of hydrogen gas should be produced in the above reaction?

Calculate your % error if in a lab situation you did the above reaction and collected 19.8 L of hydrogen gas.