

(pgs 464-468)

At sea level what percentage of air is nitrogen _____, oxygen _____
Carbon dioxide _____

What causes the “bends”

Write down Dalton’s law of partial pressures in words:

And as a formula,

Using Dalton’s law of partial pressures, what would the partial pressure of oxygen be at sea level on a day when the pressure of the atmosphere is 1.00 atm.

How about for nitrogen?

How about for carbon dioxide?

A common way to collect gases in the lab is to bubble the gas through water as described in the book.
Using chart 13.2 on page 467

What is the vapor pressure of water at 20.0 Celsius?

If 3.0 liters of a oxygen gas was collected at pressure of 760 torr at 20.0 Celsius, What would the partial pressure of the oxygen gas be ?

Using the ideal gas law, calculate how many grams of oxygen gas must be present in the above question?

